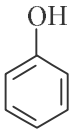
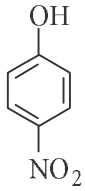
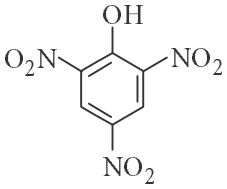
Chapter 25. Alcohols, Phenols and Ethers

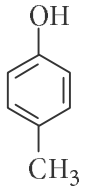
1. The heating of phenyl methyl ether with HI produces
(a) iodobenzene (b) phenol
(c) benzene (d) ethyl chloride.
(NEET 2017)

2. Which one is the most acidic compound?

(a) 

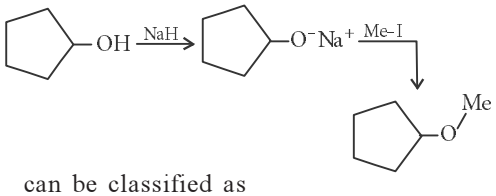
(b) 

(c) 

(d) 

(NEET 2017)

3. The reaction



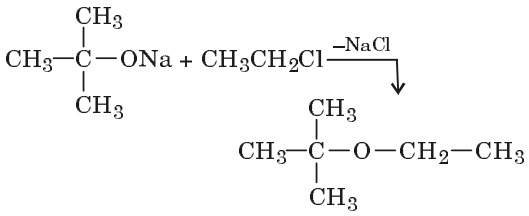
can be classified as
(a) dehydration reaction
(b) Williamson alcohol synthesis reaction
(c) Williamson ether synthesis reaction
(d) alcohol formation reaction.
(NEET-I 2016)

4. Reaction of phenol with chloroform in presence of dilute sodium hydroxide finally introduces which one of the following functional group?
(a) –COOH (b) $-CHCl_2$
(c) –CHO (d) $-CH_2Cl$ *(2015)*

5. Which of the following reaction(s) can be used for the preparation of alkyl halides?

(I) $CH_3CH_2OH + HCl \xrightarrow{Anh.\ ZnCl_2}$
(II) $CH_3CH_2OH + HCl \longrightarrow$
(III) $(CH_3)_3COH + HCl \longrightarrow$
(IV) $(CH_3)_2CHOH + HCl \xrightarrow{Anh.\ ZnCl_2}$
(a) (I) and (II) only
(b) (IV) only
(c) (III) and (IV) only
(d) (I), (III) and (IV) only *(2015)*

6. The reaction,

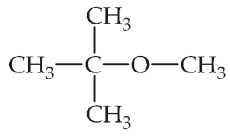


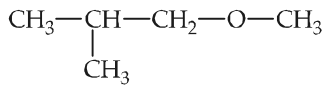
is called
(a) Etard reaction
(b) Gattermann-Koch reaction
(c) Williamson synthesis
(d) Williamson continuous etherification process. *(2015, Cancelled)*

7. Among the following sets of reactants which one produces anisole?
(a) CH_3CHO ; $RMgX$
(b) C_6H_5OH ; NaOH ; CH_3I
(c) C_6H_5OH ; neutral $FeCl_3$
(d) $C_6H_5CH_3$; CH_3COCl ; $AlCl_3$ *(2014)*

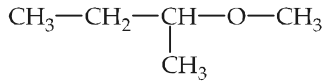
8. Which of the following will not be soluble in sodium hydrogen carbonate?
(a) 2,4,6-Trinitrophenol
(b) Benzoic acid
(c) *o*-Nitrophenol
(d) Benzenesulphonic acid *(2014)*

9. Among the following ethers, which one will produce methyl alcohol on treatment with hot concentrated HI?

(a) 

(b) 

(c) $CH_3—CH_2—CH_2—O—CH_3$

(d) 

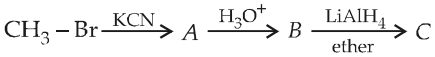
(NEET 2013)

10. Number of isomeric alcohols of molecular formula $C_6H_{14}O$ which give positive iodoform test is

(a) three (b) four
(c) five (d) two

(Karnataka NEET 2013)

11. In the following sequence of reactions



the end product (*C*) is

(a) acetone (b) methane
(c) acetaldehyde (d) ethyl alcohol

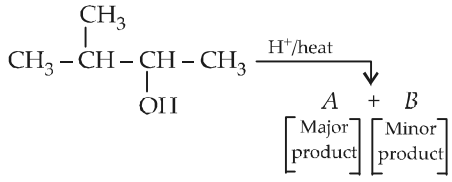
(2012)

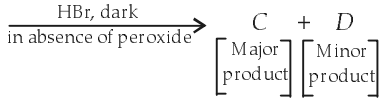
12. Which of the following compounds can be used as antifreeze in automobile radiators?

(a) Methyl alcohol (b) Glycol
(c) Nitrophenol (d) Ethyl alcohol

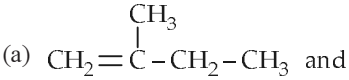
(Mains 2012)

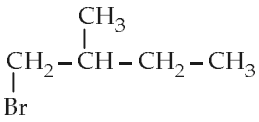
13. In the following reactions,

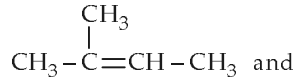
(i) 

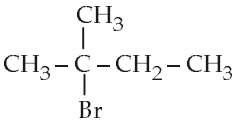
(ii) 

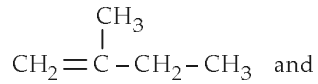
the major products (*A*) and (*C*) are respectively

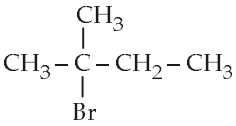
(a)  and

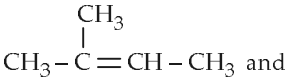


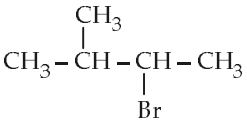
(b)  and



(c)  and



(d)  and



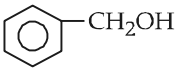
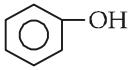
(2011)

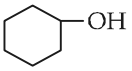
14. Given are cyclohexanol (I), acetic acid (II), 2,4,6-trinitrophenol (III) and phenol (IV). In these the order of decreasing acidic character will be

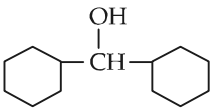
(a) III > II > IV > I (b) II > III > I > IV
(c) II > III > IV > I (d) III > IV > II > I

(2010)

15. Which of the following compounds has the most acidic nature?

(a)  (b) 

(c) 

(d) 

(2010)

16. Among the following four compounds

(i) Phenol
(ii) Methyl phenol
(iii) *Meta*-nitrophenol
(iv) *Para*-nitrophenol

The acidity order is
(a) (iv) > (iii) > (i) > (ii)
(b) (iii) > (iv) > (i) > (ii)
(c) (i) > (iv) > (iii) > (ii)
(d) (ii) > (i) > (iii) > (iv) *(2010)*

17. When glycerol is treated with excess of HI, it produces
(a) 2-iodopropane (b) allyl iodide
(c) propene (d) glycerol triiodide *(Mains 2010)*

18. Following compounds are given
(i) CH_3CH_2OH (ii) CH_3COCH_3
(iii) $CH_3-\underset{\displaystyle CH_3}{\underset{|}{CHOH}}$ (iv) CH_3OH

Which of the above compound(s), on being warmed with iodine solution and NaOH, will give iodoform?
(a) (i), (iii) and (iv) (b) Only (ii)
(c) (i), (ii) and (iii) (d) (i) and (ii) *(Mains 2010)*

19. Match the compounds given in List I with their characteristic reactions given in List II. Select the correct option.

List I (Compounds)	List II (Reactions)
A. $CH_3(CH_2)_3NH_2$	(i) Alkaline hydrolysis
B. $CH_3C{\equiv}CH$	(ii) With KOH (alcohol) and $CHCl_3$ produces bad smell
C. $CH_3CH_2COOCH_3$	(iii) Gives white ppt. with ammoniacal $AgNO_3$
D. $CH_3CH(OH)CH_3$	(iv) With Lucas reagent cloudiness appears after 5 minutes

(a) A-(ii), B-(i), C-(iv), D-(iii)
(b) A-(iii), B-(ii), C-(i), D-(iv)
(c) A-(ii), B-(iii), C-(i), D-(iv)
(d) A-(iv), B-(ii), C-(iii), D-(i) *(Mains 2010)*

20. Consider the following reaction:

$$\text{Ethanol} \xrightarrow{PBr_3} X \xrightarrow{\text{alc. KOH}} Y \xrightarrow[\text{(ii) } H_2O,\ \text{heat}]{\text{(i) } H_2SO_4,\ \text{room temperature}} Z$$

the product Z is
(a) $CH_3CH_2-O-CH_2-CH_3$
(b) $CH_3-CH_2-O-SO_3H$
(c) CH_3CH_2OH
(d) $CH_2{=}{=}CH_2$ *(2009)*

21. $HOCH_2{\cdot}CH_2OH$ on heating with periodic acid gives
(a) 2HCOOH (b) $\begin{matrix}CHO\\|\\CHO\end{matrix}$
(c) $2\begin{matrix}H\\H\end{matrix}\!\!>\!C{=}O$ (d) $2CO_2$ *(2009)*

22. Consider the following reaction:

$$\text{Phenol} \xrightarrow{\text{Zn dust}} X \xrightarrow[\text{Anhyd. } AlCl_3]{CH_3Cl} Y \xrightarrow{\text{Alkaline } KMnO_4} Z$$

the product Z is
(a) benzaldehyde (b) benzoic acid
(c) benzene (d) toluene *(2009)*

23. In the reaction:

$$CH_3-\underset{\displaystyle CH_3}{\underset{|}{CH}}-CH_2-O-CH_2-CH_3 + HI \xrightarrow{\text{heated}}$$

Which of the following compounds will be formed?

(a) $CH_3-\underset{\displaystyle CH_3}{\underset{|}{CH}}-CH_3 + CH_3CH_2OH$

(b) $CH_3-\underset{\displaystyle CH_3}{\underset{|}{CH}}-CH_2OH + CH_3CH_3$

(c) $CH_3-\underset{\displaystyle CH_3}{\underset{|}{CH}}-CH_2OH + CH_3CH_2I$

(d) $CH_3-\underset{\displaystyle CH_3}{\underset{|}{CH}}-CH_2-I + CH_3CH_2OH$

(2007)

24. The general molecular formula, which represents the homologous series of alkanols is
(a) $C_nH_{2n+2}O$ (b) $C_nH_{2n}O_2$
(c) $C_nH_{2n}O$ (d) $C_nH_{2n+1}O$ *(2006)*

25. Ethylene oxide when treated with Grignard reagent yields

(a) primary alcohol
(b) secondary alcohol
(c) tertiary alcohol
(d) cyclopropyl alcohol. *(2006)*

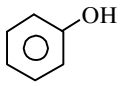
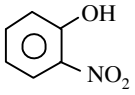
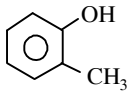
26. The major organic product in the reaction is

$CH_3 - O - CH(CH_3)_2 + HI \rightarrow$ products

(a) $CH_3I + (CH_3)_2CHOH$
(b) $CH_3OH + (CH_3)_2CHI$
(c) $ICH_2OCH(CH_3)_2$ (d) $CH_3OC(CH_3)_2$ with I below

(2006)

27. Which one of the following compounds is most acidic?

(a) $Cl - CH_2 - CH_2 - OH$
(b) 
(c) 
(d) 

(2005)

28. Which one of the following will not form a yellow precipitate on heating with an alkaline solution of iodine?

(a) $CH_3CH(OH)CH_3$
(b) $CH_3CH_2CH(OH)CH_3$
(c) CH_3OH
(d) CH_3CH_2OH *(2004)*

29. *n*-propyl alcohol and isopropyl alcohol can be chemically distinguished by which reagent

(a) PCl_5 (b) reduction
(c) oxidation with potassium dichromate
(d) ozonolysis. *(2002)*

30. When phenol is treated with $CHCl_3$ and NaOH, the product formed is

(a) benzaldehyde (b) salicylaldehyde
(c) salicylic acid (d) benzoic acid.

(2002)

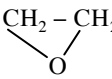
31. Which of the following is correct?

(a) On reduction, any aldehyde gives secondary alcohol.
(b) Reaction of vegetable oil with H_2SO_4 gives glycerine.
(c) Alcoholic iodine with NaOH gives iodoform.
(d) Sucrose on reaction with NaCl gives invert sugar. *(2001)*

32. Iodoform test is not given by

(a) ethanal (b) ethanol
(c) 2-pentanone (d) 3-pentanone

(1998)

33. Reaction of  with *R*Mg*X* leads to the formation of

(a) RCH_2CH_2OH (b) $RCHOHCH_3$
(c) *R*CHOHR (d) 

(1998)

34. Which one of the following compounds is resistant to nucleophilic attack by hydroxyl ions?

(a) Diethyl ether (b) Acetonitrile
(c) Acetamide (d) Methyl acetate

(1998)

35. When 3,3-dimethyl-2-butanol is heated with H_2SO_4, the major product obtained is

(a) 2,3-dimethyl-2-butene
(b) *cis* and *trans* isomers of 2,3-dimethyl-2-butene
(c) 2,3-dimethyl-1-butene
(d) 3,3-dimethyl-1-butene. *(1995)*

36. On heating glycerol with conc. H_2SO_4, a compound is obtained which has bad odour. The compound is

(a) acrolein (b) formic acid
(c) allyl alcohol
(d) glycerol sulphate. *(1994)*

37. The compound which does not react with sodium is

(a) CH_3COOH (b) $CH_3CHOHCH_3$
(c) C_2H_5OH (d) CH_3OCH_3

(1994)

38. Ethanol and dimethyl ether form a pair of functional isomers. The boiling point of ethanol is higher than that of dimethyl ether, due to the presence of

(a) H-bonding in ethanol
(b) H-bonding in dimethyl ether
(c) CH_3 group in ethanol
(d) CH_3 group in dimethyl ether. *(1993)*

39. Increasing order of acid strength among *p*-methoxyphenol, *p*-methylphenol and *p*-nitrophenol is

(a) *p*-nitrophenol, *p*-methoxyphenol, *p*-methylphenol

(b) *p*-methylphenol, *p*-methoxyphenol, *p*-nitrophenol
(c) *p*-nitrophenol, *p*-methylphenol, *p*-methoxyphenol
(d) *p*-methoxyphenol, *p*-methylphenol, *p*-nitrophenol. *(1993)*

40. Which one of the following on oxidation gives a ketone?
(a) Primary alcohol
(b) Secondary alcohol
(c) Tertiary alcohol
(d) All of these. *(1993)*

41. What is formed when a primary alcohol undergoes catalytic dehydrogenation?
(a) Aldehyde (b) Ketone
(c) Alkene (d) Acid *(1993)*

42. When hydrochloric acid gas is treated with propene in presence of benzoyl peroxide, it gives
(a) 2-chloropropane
(b) allyl chloride
(c) no reaction
(d) *n*-propyl chloride. *(1993)*

43. How many isomers of $C_5H_{11}OH$ will be primary alcohols?
(a) 5 (b) 4
(c) 2 (d) 3 *(1992)*

44. Methanol is industrially prepared by
(a) oxidation of CH_4 by steam at 900°C
(b) reduction of HCHO using $LiAlH_4$
(c) reaction HCHO with a solution of NaOH
(d) reduction of CO using H_2 and $ZnO–Cr_2O_3$. *(1992)*

45. HBr reacts fastest with
(a) 2-Methylpropan-1-ol
(b) Methylpropan-2-ol
(c) propan-2-ol
(d) propan-1-ol. *(1992)*

46. When phenol is treated with excess bromine water. It gives
(a) *m*-bromophenol
(b) *o*-and *p*-bromophenols
(c) 2,4-dibromophenol
(d) 2,4,6-tribromophenol. *(1992)*

47. The compound which reacts fastest with Lucas reagent at room temperature is
(a) butan-1-ol
(b) butan-2-ol
(c) 2-methylpropan-1-ol
(d) 2-methylpropan-2-ol. *(1989)*

48. Which one of the following compounds will be most readily attacked by an electrophile?
(a) Chlorobenzene (b) Benzene
(c) Phenol (d) Toluene *(1989)*

49. Propene, $CH_3CH{=}CH_2$ can be converted into 1-propanol by oxidation. Indicate which set of reagents amongst the following is ideal for the above conversion?
(a) $KMnO_4$ (alkaline)
(b) Osmium tetroxide (OsO_4/CH_2Cl_2)
(c) B_2H_6 and alk. H_2O_2
(d) O_3/Zn. *(1989)*

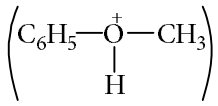
50. Phenol is heated with $CHCl_3$ and aqueous KOH when salicylaldehyde is produced. This reaction is known as
(a) Rosenmund's reaction
(b) Reimer-Tiemann reaction
(c) Friedel-Crafts reaction
(d) Sommelet reaction. *(1989,88)*

51. Lucas reagent is
(a) conc. HCl and anhydrous $ZnCl_2$
(b) conc. HNO_3 and hydrous $ZnCl_2$
(c) conc. HCl and hydrous $ZnCl_2$
(d) conc. HNO_3 and anhydrous $ZnCl_2$. *(1988)*

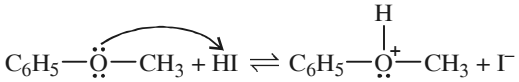
Answer Key

1.	(b)	**2.**	(c)	**3.**	(c)	**4.**	(c)	**5.**	(d)	**6.**	(c)	**7.**	(b)	**8.**	(c)	**9.**	(a)	**10.**	(b)
11.	(d)	**12.**	(b)	**13.**	(b)	**14.**	(a)	**15.**	(b)	**16.**	(a)	**17.**	(a)	**18.**	(c)	**19.**	(c)	**20.**	(c)
21.	(c)	**22.**	(b)	**23.**	(c)	**24.**	(a)	**25.**	(a)	**26.**	(a)	**27.**	(c)	**28.**	(c)	**29.**	(c)	**30.**	(b)
31.	(c)	**32.**	(d)	**33.**	(a)	**34.**	(a)	**35.**	(a)	**36.**	(a)	**37.**	(d)	**38.**	(a)	**39.**	(d)	**40.**	(b)
41.	(a)	**42.**	(a)	**43.**	(b)	**44.**	(d)	**45.**	(b)	**46.**	(d)	**47.**	(d)	**48.**	(c)	**49.**	(c)	**50.**	(b)
51.	(a)																		

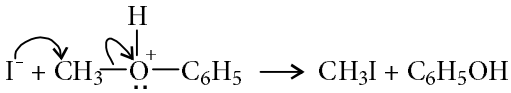
EXPLANATIONS

1. (b) : In case of phenyl methyl ether, methyl phenyl oxonium ion  is formed by protonation of ether. The O—CH_3 bond is weaker than O—C_6H_5 bond as O—C_6H_5 has partial double bond character. Therefore, the attack by I^- ion breaks O—CH_3 bond to form CH_3I.

Step I :



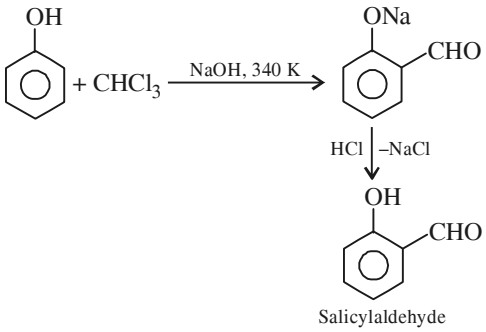
Step II :



2. (c) : Electron withdrawing groups increase the acidity while electron donating groups decrease the acidity of phenol.

3. (c) : Williamson's ether synthesis reaction involves the treatment of sodium alkoxide with a suitable alkyl halide to form an ether.

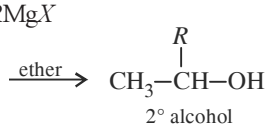
4. (c) : This is Reimer–Tiemann reaction.

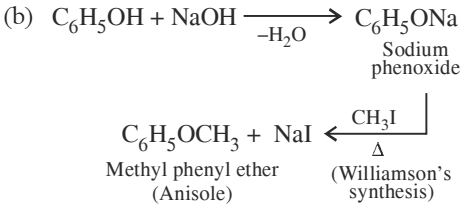


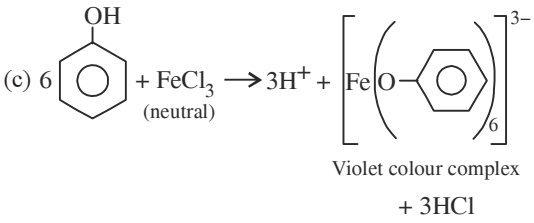
5. (d) : 1° and 2° alcohols react with HCl in presence of anhydrous $ZnCl_2$ as catalyst while in case of 3° alcohols $ZnCl_2$ is not required.

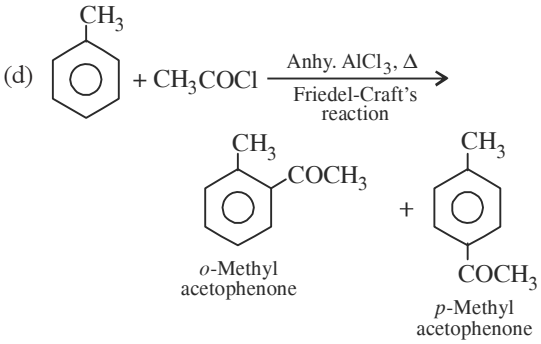
6. (c) : Williamson synthesis is the best method for the preparation of ethers.

7. (b) : (a) $CH_3CHO + RMgX$



(b) 

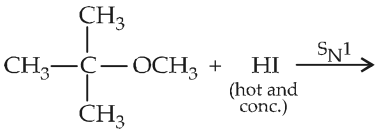
(c) 

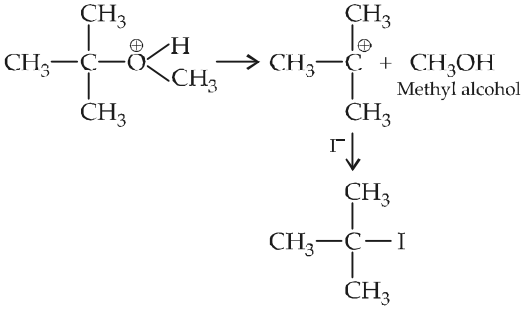
(d) 

8. (c) : The reaction is as follows:

Acid + $NaHCO_3$ ⟶ Sodium salt of acid (soluble) + H_2CO_3

Among all the given compounds, *o*-nitrophenol is weaker acid than HCO_3^-. Hence, it does not react with $NaHCO_3$.

9. (a) : 



10. (b): The iodoform test is positive for alcohols with formula R — CHOH — CH_3. Among $C_6H_{14}O$ isomers, the ones with positive iodoform test are:

I. $CH_3 — CH_2 — CH_2 — CH_2 — CHOH — CH_3$
2- Hexanol

II. $CH_3 — CH_2 — CH(CH_3) — CHOH — CH_3$
3- Methyl-2-Pentanol

III. $(CH_3)_2CH — CH_2 — CHOH — CH_3$
4- Methyl-2-Pentanol

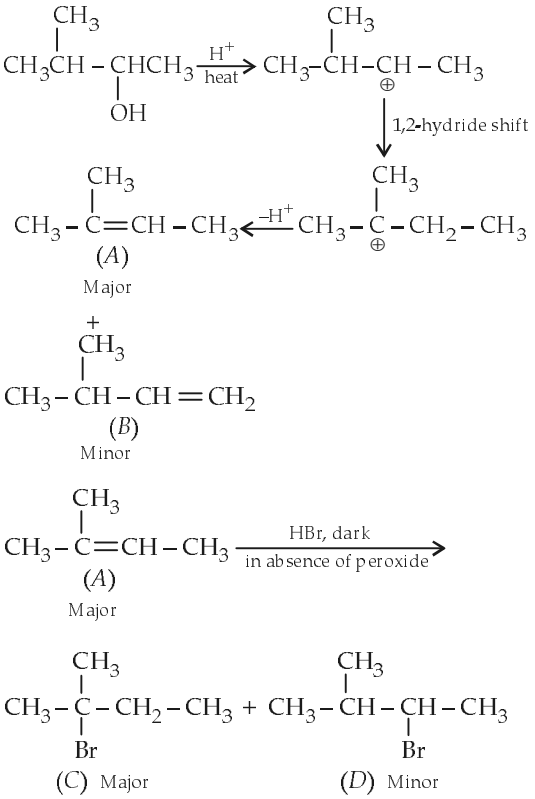
IV. $(CH_3)_3C — CHOH — CH_3$
3, 3- Dimethyl-2-butanol

11. (d): $CH_3Br \xrightarrow{KCN} \underset{(A)}{CH_3CN} \xrightarrow{H_3O^+}$

$\underset{(B)}{CH_3COOH} \xrightarrow[ether]{LiAlH_4} \underset{(C)}{CH_3CH_2OH}$

12. (b)

13. (b):



14. (a): III > II > IV > I

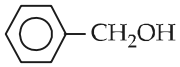
Since, phenols and carboxylic acids are more acidic than aliphatic alcohols, we find that cyclohexanol (I) is least acidic. Out of the two given phenols, III is more acidic than IV. This is because of the presence of three highly electron withdrawing $–NO_2$ groups on the benzene ring which makes the O—H bond extremely polarized. This facilitates the release of H as H^+. Thus, III > IV.

In acetic acid the electron withdrawing $—\overset{\overset{O}{\|}}{C}—$ in the —COOH group polarises the O—H bond and increases the acidic strength. Acetic acid is therefore more acidic than phenol or cyclohexanol.

∴ The order is III > II > IV > I.

15. (b): Phenol is most acidic of all the given compounds.

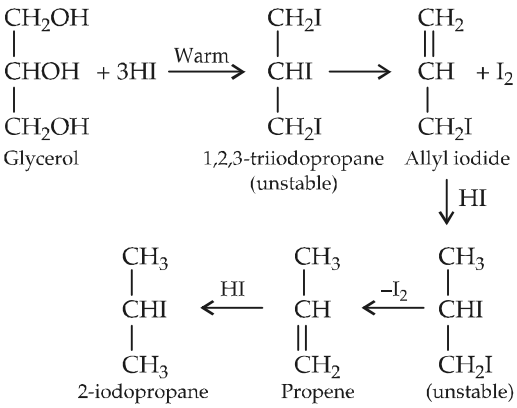
In phenol, the electron withdrawing phenyl ring polarizes the O—H bond thereby facilitating the release of H as H^+ and hence phenol is most acidic.

In  the electron withdrawing effect of phenyl ring is somewhat diminished by the $–CH_2$ group and it is therefore less acidic than phenol. In (c) and (d), —OH group is attached to alkyl groups which, due to their $+I$ effect reduce the polarity of —OH bond and so the acidic strength is low.

16. (a): In phenols, the presence of electron releasing groups decrease the acidity, whereas presence of electron withdrawing groups increase the acidity, compared to phenol. Among the *meta* and *para*-nitrophenols, the latter is more acidic as the presence of $–NO_2$ group at *para* position stabilises the phenoxide ion to a greater extent than when it is present at *meta* position. Thus, correct order of acidity is :

Para-nitrophenol > *meta*-nitrophenol > phenol > methyl phenol
(iv) (iii) (i) (ii)

17. (a):



18. (c) : Methyl alcohol does not respond to the iodoform test. The iodoform test is exhibited by ethyl alcohol, acetaldehyde, acetone, methyl ketones, those alcohols which possess $CH_3CH(OH)$— group, acetophenone, α-hydroxypropionic acid, keto acid, 2-aminoalkanes, etc.

19. (c) : $\underset{1° \text{Amine}}{CH_3CH_2CH_2CH_2NH_2} \xrightarrow[CHCl_3]{KOH\ (alc.)}$

$CH_3CH_2CH_2CH_2NC + 3KCl + 3H_2O$
Bad smell

$CH_3C{\equiv}CH + Amm.\ AgNO_3 \longrightarrow$

$CH_3C{\equiv}\bar{C}Ag^+ + HNO_3$
White ppt.

$CH_3CH_2COOCH_3 + NaOH \xrightarrow{Heat}$
$CH_3CH_2COONa + CH_3OH$

$\underset{2° \text{ Alcohol}}{CH_3CH(OH)CH_3} + \text{Conc. } HCl \xrightarrow{anhy.\ ZnCl_2}$

$CH_3-\underset{\underset{Cl}{|}}{CH}-CH_3 + H_2O$
cloudiness
appears in 5 minutes

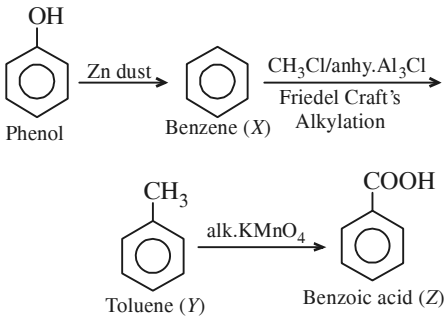
20. (c) :

$\underset{\text{Ethanol}}{C_2H_5OH} \xrightarrow{PBr_3} \underset{(X)}{C_2H_5Br} \xrightarrow[\text{dehydro-halogenation}]{\text{alc. KOH}}$

$\underset{\text{Ethene}}{\underset{(Y)}{CH_2=CH_2}} \xrightarrow{H_2SO_4} CH_3CH_2-HSO_4 \xrightarrow{H_2O} \underset{(Z)}{C_2H_5-OH}$

21. (c) : When 1,2-diol like ethylene glycol is treated with HIO_4, each alcoholic group is oxidised to a carbonyl group by HIO_4. Since in glycol, both the –OH groups are terminal, so oxidation would yield two formaldehyde molecules.

i.e., $\begin{matrix}CH_2-OH\\|\\CH_2-OH\end{matrix} \xrightarrow{HIO_4} 2HCHO$

22. (b) : 

23. (c) : The alkyl iodide produced depends on the nature of the alkyl groups. If one group is Me and the other a primary or secondary alkyl group, it is methyl iodide which is produced. This can be explained on the assumption that the mechanism is S_N2, and because of the steric effect of the larger group, I^- attacks the smaller methyl group.
When the substrate is a methyl *t*-alkyl ether, the products are *t*-*R*I and MeOH. This can be explained by an S_N1 mechanism, the carbonium ion produced being the *t*-*R* since tertiary carbonium ion is more stable than a primary or secondary carbonium ion.

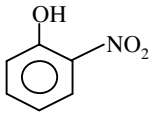
24. (a) : All alcohols follow the general formula $C_nH_{2n+2}O$.
$CH_3OH\ [CH_{2+2}O]$; $C_2H_5OH\ [C_2H_{(2\times2)+2}O]$
$n = 1$, $n = 2$

25. (a) : $H_2C\overset{O}{\diagup\diagdown}CH_2 + CH_3MgBr \longrightarrow$

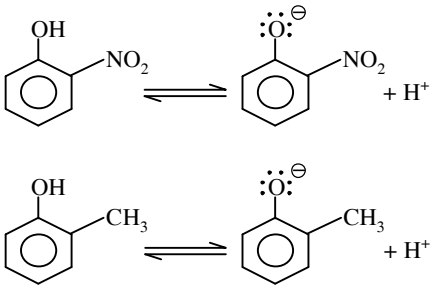
$CH_3CH_2CH_2OMgBr \xrightarrow{H^+} \underset{\text{primary alcohol}}{CH_3CH_2CH_2OH}$

26. (a) : With cold HI, a mixture of alkyl iodide and alcohol is formed. In the case of mixed ethers, the halogen atom attaches to a smaller and less complex alkyl group.
$CH_3OCH(CH_3)_2 + HI \rightarrow CH_3I + (CH_3)_2CHOH$

27. (c) : Phenols are much more acidic than alcohols, due to the stabilisation of phenoxide ion by resonance.



—NO_2 is the electron withdrawing group and helps in stabilizing the negative charge on the oxygen hence equilibrium shifts in forward direction and more H^+ remove easily. Hence it is most acidic.



—CH_3 is the electron donating group.
Hence electron density increases on the oxygen and destabilizes the product. Thus, equilibrium shifts in backward direction.

28. (c) : Formation of a yellow precipitate on heating a compound with an alkaline solution of iodine is known as iodoform reaction. Methyl alcohol does not respond to this test. Iodoform test is exhibited by ethyl alcohol, acetaldehyde, acetone, methyl ketones and those alcohols which possess $CH_3CH(OH)$ – group.

29. (c) : *n*-propyl alcohol on oxidation with potassium dichromate gives an aldehyde which on further oxidation gives an acid. Both aldehyde and acid contain the same number of C atoms as the original alcohol.

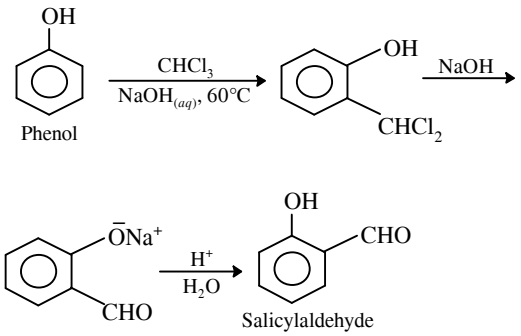
$$CH_3CH_2CH_2OH \xrightarrow[K_2Cr_2O_7/H_2SO_4]{25^\circ C}$$

$$CH_3CH_2CHO \xrightarrow{K_2Cr_2O_7/H_2SO_4} CH_3CH_2COOH$$

Isopropyl alcohol on oxidation gives a ketone with the same number of C atoms as the original alcohol.

$$CH_3-\underset{\displaystyle CH_3}{\overset{\displaystyle OH}{\underset{|}{\overset{|}{CH}}}} \xrightarrow{K_2Cr_2O_7/H_2SO_4} CH_3-\overset{\displaystyle O}{\overset{||}{C}}-CH_3$$

30. (b) : This reaction is called Reimer-Tiemann reaction.



31. (c) : $C_2H_5OH + 4I_2 + NaOH \rightarrow CHI_3 + NaI + HCOONa + H_2O$

Iodoform is a pale yellow solid which crystallises in hexagonal plates.

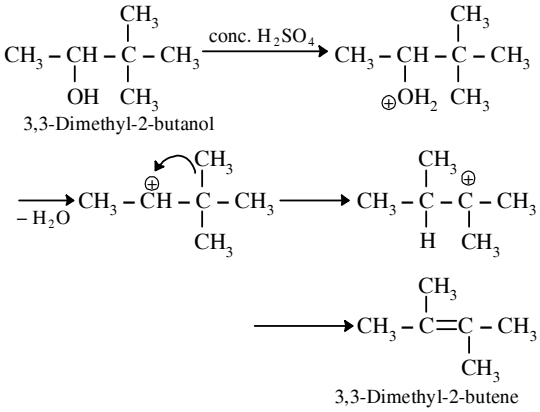
32. (d) : Ethyl alcohol, 2-alkanols, and carbonyl compounds containing $CH_3-\overset{\displaystyle O}{\overset{||}{C}}-$ group show iodoform test, *i.e.*, acetaldehyde and 2-ketones etc. So iodoform test is not given by 3-pentanone.

33. (a) :

$$R-Mg-X + CH_2-\underset{\displaystyle O}{\underset{\diagdown\ \diagup}{}}CH_2 \xrightarrow[\text{Ethylene oxide}]{} \underset{\text{Adduct}}{R-CH_2-CH_2-OMgX} \xrightarrow[\Delta]{H_2O/H^+} \underset{1^\circ\text{-alcohol}}{R-CH_2-CH_2-OH} + Mg(OH)X$$

34. (a) : Diethyl ether is a saturated compound, so it is resistant to nucleophilic attack by a hydroxyl ion (OH^-).
Other compounds have unsaturation and the unsaturated 'C' atom bears partial +ve charge, therefore they undergo easy nucleophilic attack by OH^- ion.

35. (a) :



36. (a) :

$$\underset{\text{Glycerol}}{\begin{matrix}CH_2OH\\|\\CHOH\\|\\CH_2OH\end{matrix}} \xrightarrow{\text{conc. } H_2SO_4} \underset{\text{Acrolein}}{CH_2=CHCHO} + 2H_2O$$

37. (d) : Ethers are very inert. The chemical inertness of ethers is due to absence of active group

in their molecules. Since $CH_3 - O - CH_3$ is an inert and it does not contain active group, therefore it does not react with sodium.

38. (a) : Hydrogen bonding in ethanol.

39. (d) : $-OCH_3$, $-CH_3$ are electron donating groups decreases the acidic character of phenols. $-NO_2$, $-CN$ are electron withdrawing groups, tend to increase the acidic character. Thus, the order is *p*-methoxyphenol < *p*-methylphenol < *p*-nitrophenol.

40. (b) : 2° alcohols on oxidation give ketones, 1° alcohols form aldehydes.

41. (a) : Primary alcohol undergoes catalytic dehydrogenation to give aldehyde.

42. (a) : Addition of HCl to propene even in the presence of benzoyl peroxide occurs according to Markownikov's rule.

$$CH_3-CH=CH_2 \xrightarrow[(C_6H_5COO)_2]{HCl} \underset{\text{2-chloropropane}}{CH_3-\underset{Cl}{CH}-CH_3}$$

43. (b) : 4-isomers are possible for $C_5H_{11}OH$.

(i) $CH_3CH_2CH_2CH_2CH_2OH$

(ii) $CH_3CH_2-\underset{CH_3}{\underset{|}{CH}}-CH_2OH$

(iii) $CH_3-\underset{CH_3}{\underset{|}{CH}}-CH_2-CH_2OH$

(iv) $CH_3-\underset{CH_3}{\overset{CH_3}{\overset{|}{\underset{|}{C}}}}-CH_2OH$

44. (d) : $CO + 2H_2 \xrightarrow{ZnO - Cr_2O_3} CH_3OH$

45. (b) : $CH_3-\overset{CH_3}{\overset{|}{\underset{OH}{\underset{|}{C}}}}-CH_3$, generates 3° carbocation which is very stable intermediate, thus it will react more rapidly with HBr.

46. (d) : Phenol on reaction with excess bromine water gives 2,4,6-tribromophenol.

47. (d) : 2-Methylpropan-2-ol reacts rapidly with Lucas reagent at room temperature.

$$\underset{\text{2-Methyl-2-propanol}}{CH_3-\overset{CH_3}{\overset{|}{\underset{CH_3}{\underset{|}{C}}}}-OH} \xrightarrow{H^+} CH_3-\overset{CH_3}{\overset{|}{\underset{CH_3}{\underset{|}{C^{\oplus}}}}} \xrightarrow{Br^-} CH_3-\overset{CH_3}{\overset{|}{\underset{CH_3}{\underset{|}{C}}}}-Br$$

48. (c) : –OH group being electron donor increases the electron density in phenol. Thus, the electron density in phenol is higher than that of toluene, benzene and chlorobenzene.

49. (c) :

$$CH_3-CH=CH_2 \xrightarrow[\text{Hydroboration}]{BH_3,\ THF} \underset{\text{Tripropyl borane}}{(CH_3CH_2CH_2)_3B} \xrightarrow[OH^-]{H_2O_2} \underset{\text{Propan-1-ol}}{3CH_3CH_2CH_2OH}$$

50. (b) : Treatment of phenol with $CHCl_3$ and aqueous hydroxide introduces – CHO group, onto the aromatic ring generally *ortho* to the – OH group. This reaction is known as Reimer - Tiemann reaction.

51. (a) : Conc. HCl and anhydrous $ZnCl_2$–Lucas reagent.